

## Reactions Remediation

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How did you do by standard:

#3-1 \_\_\_\_/12 Can I translate formulas into chemical reactions.

#3-2: \_\_\_\_/3 I can represent ions in solution proportionally and accurately

#5-7 \_\_\_\_/3 I can predict the products of a solubility reaction and correctly predict the state of the products.

#5-8 \_\_\_\_/9 I can predict the products of a oxidation reduction reaction, and correctly calculate the voltages?

#5-9 \_\_\_\_/1 I can utilize the factors affecting the rate to speed up or slow down a reaction.

#3: To translate names of chemicals into formulas you must know how a substance is bonding. The type of bond will identify the procedure used to name the substance. Whether that be ionic, covalent, or acidic.

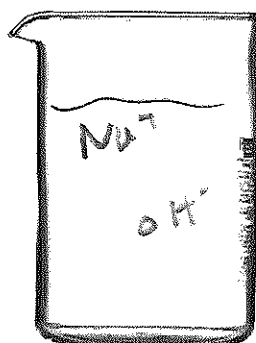
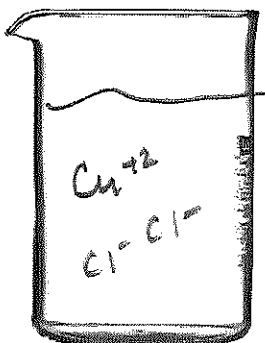


1. Convert the following names to formulas and vice versa.

- Hydronitric acid =  $H_3N(aq)$        $NH_4NO_3 =$  Ammonium Nitrate
- Aluminum chloride =  $NH_4Cl$        $CaCl_2 =$  Calcium Chloride
- Sulfur trioxide =  $SO_2$        $BF_3 =$  Boron tri fluoride
- Copper (II) sulfate =  $CuSO_4$        $H_2SO_3(aq) =$  Sulfurous Acid
- Nitrous Acid =  $HNO_2(aq)$        $HF(aq) =$  Hydrofluoric Acid

#3 Particulate representation of ions in solutions. Go to time 11:48 on video.

2. In the beaker below draw a particulate model of the following two equal concentration solutions.  $CuCl_2$  and  $NaOH$

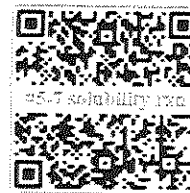
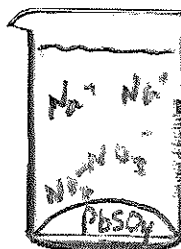
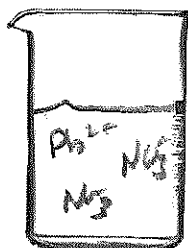
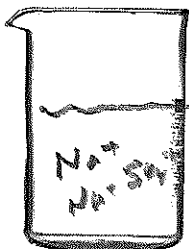


#5-7: Solubility based chemical reactions

3. 0.2M Sodium sulfate is poured into 0.2M lead (II) nitrate solution. Translate and determine the products of the reaction.

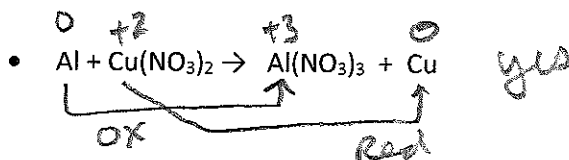
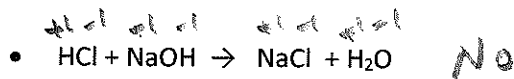


4. Draw both beakers proportionally and the products proportionally in the 3 beakers provided.

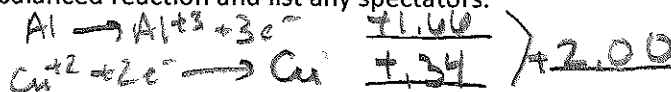


#5-8: Oxidation Reactions.

5. With the following two reactions indicate if the reaction is an oxidation reduction reaction and if so, who is being oxidized and reduced.



6. For any reaction above that is a redox reaction, write out the half reactions, calculate the resulting voltage, write the balanced reaction and list any spectators.



7. Draw out a particulate diagram of the redox reaction.

8. Copper when dropped into HCl does not cause a reaction. Justify this with calculations.

